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## Microwave-Assisted Synthesis of ZnO Micro-Javelins

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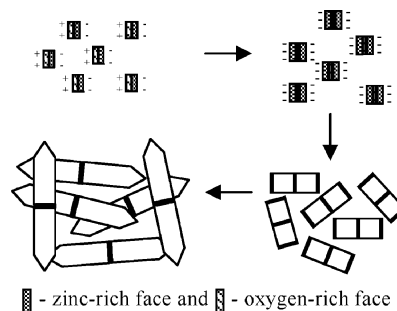
**Authors**

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### Microwave-assisted synthesis of ZnO micro-javelins

Sibu C. Padmanabhan,\* Deirdre Ledwith, Suresh C. Pillai, Declan E. McCormack and John M. Kelly\*

A rapid and straightforward microwave-assisted method for preparing well-defined double-ended ZnO microneedles (micro-javelins) from simple reagents is reported. The process produces reasonably defect-free monodisperse ZnO micro-javelins with a high yield.



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# Microwave-assisted synthesis of ZnO micro-javelins†

Sibu C. Padmanabhan,‡\*<sup>a</sup> Deirdre Ledwith,<sup>a</sup> Suresh C. Pillai,<sup>b</sup> Declan E. McCormack<sup>c</sup> and John M. Kelly\*<sup>a</sup>

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The microwave (MW)-assisted formation of ZnO micro-javelins from zinc nitrate and urea in aqueous solution is described. The particles (named as 'micro-javelins' because of their high aspect ratio and needle-like tips) grow hexagonally with well-defined facets in the (01–10) direction and pointed tips in (0001) direction. Powder X-ray diffraction patterns show the appearance of a strikingly dominant (1000) orientation. Scanning electron microscopy (SEM) and transmission electron microscopy (TEM) investigations reveal the morphological evolution of these hexagonal ZnO particles with time. The effect of precursor concentrations, counterion type and MW irradiation power and their consequent influence on pH and Zn<sup>2+</sup> ion concentration are investigated. A mechanism for the formation of the micro-javelins is postulated. The microwave induced supersaturation of Zn(OH)<sup>+</sup> species under the weakly basic pH condition and the initial growth through the (000–1) direction (oxygen-rich face) are proposed to be the key factors that dictate the formation of these ZnO micro-javelins. The present one-step microwave process is a straightforward and a reproducible method for the bulk synthesis of defect-free ZnO micro-javelins, which would find potential applications in microelectronic devices (*e.g.* lasers, cantilevers in surface probing equipment, *etc.*).

## Introduction

ZnO, a wide band gap semiconductor (3.37 eV), has a range of potential applications as photocatalysts,<sup>1</sup> gas sensors,<sup>2</sup> lasers,<sup>3</sup> displays,<sup>4</sup> high-density optical storage devices<sup>5</sup> and varistors.<sup>6</sup> Recent research has focused on manipulating its shape into one-dimensional (1-D) rods,<sup>7</sup> tubes,<sup>8</sup> belts,<sup>9</sup> wires and rings<sup>10</sup> that offer improved performance in terms of its optical, mechanical and electronic properties.

Solid–vapour phase sublimation is one of the established techniques in the preparation of 1-D ZnO nanoparticles.<sup>11</sup> By this method the growth of the reactants into the required 1-D structures is particularly sensitive to reaction conditions including the selection of carrier gas, its flow rate and the orientation of substrate, which are some of its limitations. Solution-growth, in this regard, may address some of the issues of the solid–vapour technique.<sup>12–15</sup> The hydrothermal method is one of the widely used solution-growth techniques to synthesise size- and shape-controlled crystalline ZnO particles.<sup>16,17</sup> The use of various precursor combinations including different zinc salts,<sup>18</sup> hydrolysing agents (amines,<sup>19</sup> acids<sup>20</sup> and bases<sup>21</sup>) as well

as various polymeric agents<sup>22</sup> and surfactants<sup>23</sup> can be helpful in defining the growth directions by capping/modifying the basic building units.<sup>24–26</sup>

The controlled solution-growth of ZnO particles with well-defined morphologies, however, is still a challenge and the mechanism underlying such processes is unclear. The oriented/controlled growth of capped Zn<sup>2+</sup> species, microemulsions,<sup>27</sup> and micellar<sup>28</sup> building units into various size- and shape-controlled ZnO particles is interesting but complicated. In 1990 Tsuchida and Kitajima reported a more straightforward synthetic procedure for the bulk chemical synthesis of well-defined ZnO hexagonal prisms. The procedure involved the hydrolysis of a zinc nitrate solution at 95 °C in the presence of urea.<sup>29</sup> Such a procedure was straightforward in that it did not involve any further additives. However, the procedure did not yield particles of a single morphology but a mixture of needles, rods and spindles and the procedure yielded only millimolar quantities of products.

As part of a programme to develop routes to synthesise ZnO nano- and micro-materials with optimal electronic and optical properties, we have previously observed that a one-step microwave (MW)-assisted reaction between zinc nitrate and urea in aqueous media yields ZnO micro-javelins.<sup>30</sup> Microwave synthesis is chosen because of its higher reaction rates, rapid volumetric heating, higher yields of products and energy saving prospects.<sup>31,32</sup> This paper reports a systematic study of the various parameters that influence the javelin formation and proposes a mechanistic pathway by which micro-javelins are formed. The effects of MW energy and irradiation time, influence of precursors (counterions), precursor concentrations, pH and Zn<sup>2+</sup> ion concentration on the micro-javelin formation are discussed in detail. This, to the best of our knowledge, is a first systematic study report describing the MW growth of such particles. The process is rapid and reproducible compared to the conventional

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1 heating and hydrothermal routes and also yields large quantities  
(up to 90%) of ZnO micro-javelins. The difference of our method  
5 from the reported ones is its ability to produce monodispersed  
ZnO micro-javelins with a high yield.

## Materials and methods

### Materials

10  $\text{Zn}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ ,  $\text{Zn}(\text{OOCCH}_3)_2 \cdot 2\text{H}_2\text{O}$  and  $\text{ZnSO}_4$  were  
purchased from Aldrich and  $\text{CO}(\text{NH}_2)_2$  and  $\text{ZnCl}_2$  from BDH.  
The chemicals were used without further purification. Millipore  
water was used for all experiments.

### Methods

15 In a typical experiment, aqueous solutions of 40 mL each of 10  
mM zinc nitrate and 100 mM urea were mixed in a Teflon bomb  
(Savillex Digestion vessel equipped with pressure releasable  
valve; volume 120 mL, height 16 cm and outer diameter 50 mm),  
20 stirred for 10 min and then irradiated at 150 W for 24 min using  
a domestic microwave oven (850 W Samsung FE22). The sample  
was then allowed to cool to room temperature (20 h of cooling).  
The precipitate was then separated by filtration (0.45  $\mu\text{m}$   
membrane filter) and dried at 80  $^\circ\text{C}$  in an air oven (20 h) before  
25 further analyses. Samples were also collected after sudden release  
of pressure (without the cooling stage) under otherwise similar  
experimental conditions and compared. Experiments involving  
different zinc salts such as zinc acetate, zinc chloride and zinc  
sulfate have also been carried out and results compared. It  
30 should be noted that the experiments carried out in this report  
were in a different microwave oven compared to our earlier  
report, where a domestic microwave oven of 750 W power was  
used,<sup>30</sup> and hence needed slightly different MW irradiation times  
to achieve similar particles.

### Characterisation

35 The pH of the reactant mixture as well as the product solution  
was measured using a Metrohm 632 pH-meter (Switzerland).  
Soluble zinc species concentrations of the microwave treated  
solutions were determined by complexometric titration (classic  
Zn-EDTA titration). Room temperature powder XRD patterns  
40 of the samples were recorded using a Siemens D 500 X-ray  
diffractometer in the diffraction angle range  $2\theta$  10–70 $^\circ$  using  
 $\text{CuK}\alpha$  radiation. FTIR spectra of the samples in KBr pellets were  
recorded with a Perkin Elmer Spectrum One GX Spectrometer.  
Photoluminescence of the samples suspended in an ethylene  
glycol : water (2 : 1) mixture was recorded using a Varian Cary  
Eclipse Fluorescence spectrophotometer. The excitation source  
50 was a high intensity xenon lamp and the excitation wavelength  
was selected as 325 nm. The morphology of the samples was  
imaged by field emission scanning electron microscopy (Hitachi  
S-4300 FESEM) and transmission electron microscopy (JEOL  
JEM 2100 electron microscope).

## Results

55 Well-crystallised wurtzite ZnO micro-javelins have been  
prepared by a microwave induced reaction of zinc nitrate and

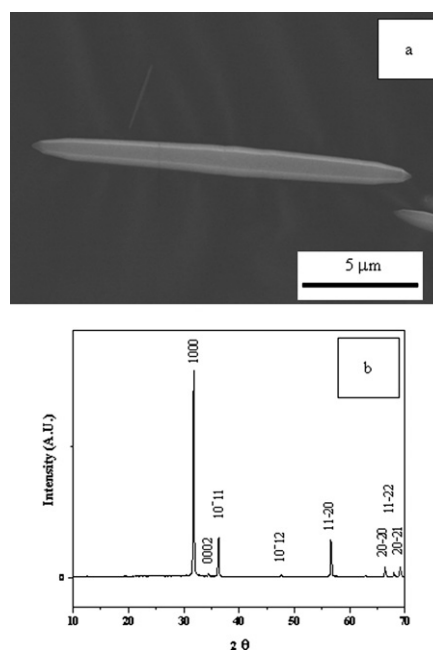


Fig. 1 (a) FESEM image and (b) XRD of ZnO micro-javelin(s) prepared by 24 min of MW irradiation between 80 mL aqueous solution of 5 mM zinc nitrate and 50 mM urea (1 : 10).

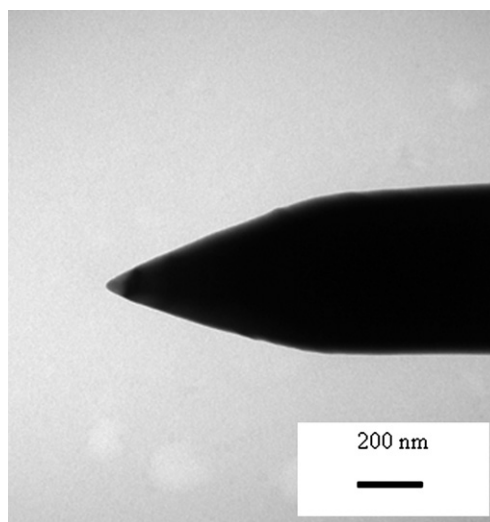


Fig. 2 TEM image of the pointed tip of a micro-javelin.

urea aqueous solution. A typical scanning electron microscope (SEM) image and X-ray diffraction (XRD) pattern of ZnO micro-javelins obtained by a 24 min of MW process between 80 mL aqueous solution of 5 mM zinc nitrate and 50 mM urea are presented in Fig. 1. The formation of well-defined micro-javelin-like ZnO particles can be observed from the SEM. The XRD pattern corresponds to hexagonal, primitive,  $P6_3mc$  wurtzite ZnO unit cell structure (ICDD card no. 36-1451). Further, the powder sample shows an unusual XRD pattern, where the (1000) peak intensity is considerably dominant compared to all other diffraction peaks. This is uncommon for most wurtzite ZnO samples, where the (1010) peak is normally dominant.

Furthermore, the extremely high intensity of the (1000) peak reflects the high crystallinity of micro-javelins. The micrographs show that the micro-javelins are grown with  $\sim 14 \mu\text{m}$  in length and  $\sim 1.1 \mu\text{m}$  in diameter and of a single morphology. Fig. 2 presents the TEM image of the pointed tip of one micro-javelin. A systematic investigation of this interesting growth phenomenon has been carried out and the results are presented here.

### Effect of precursor concentration

In order to identify the optimum reactant concentrations that yield perfectly grown ZnO micro-javelins, experiments were conducted with different concentrations of urea and zinc nitrate. A ratio of zinc nitrate to urea of 1 : 10 (5 mM : 50 mM) is found to produce a high yield of well-defined ZnO micro-javelins. No precipitate was formed when the zinc nitrate to urea ratio was 1 : 100 (5 mM : 500 mM). At a 1 : 1 zinc nitrate to urea ratio (5 mM :

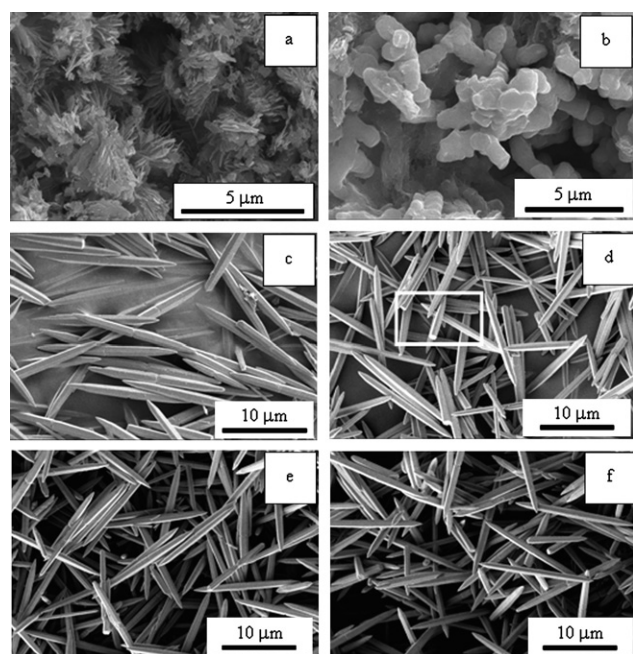
5 mM), double-ended cone-like particles were observed where the particular structure appears to be formed by the lateral stacking of disc-like particles and the particles appear to grow down to form coupled cone-like structures (ESI† 1). When the reaction was carried out with solutions containing 100 mM zinc nitrate and 100 mM urea (1 : 1), both ZnO micro-javelins and agglomerates were formed. This indicates that the zinc nitrate to urea ratio is crucial in the preferential growth of ZnO micro-javelins with a single morphology and a zinc nitrate to urea ratio of 1 : 10 produces a maximum number of micro-javelins of uniform morphology. Therefore a 1 : 10 ratio of zinc nitrate to urea was selected for further experiments to investigate the growth mechanism.

### Effect of microwave energy and time of irradiation

The effect of the period of irradiation was investigated by varying the time of reactions at a fixed power of 150 W. XRD patterns of samples obtained at different MW irradiation time were recorded (ESI† 2). 12 and 13 min irradiation produced mostly white precipitates with a mixture of  $\text{Zn}_5(\text{CO}_3)_2(\text{OH})_6$  (1) (ICDD card 19-1458) and ZnO structures. The 14 min irradiated sample on the other hand gave peaks corresponding to only wurtzite ZnO (ICDD card 36-1451). The (1000) peak showed a consistent yet remarkable rise in intensity compared to other peaks upon increasing the irradiation time further up to 24 min.

The morphological evolution of the ZnO micro-javelins was followed using SEM, which led us to some interesting observations regarding the micro-javelin formation. Despite a few reports<sup>33,34</sup> indicating the formation of such particles, no significant attempt has been made to understand the mechanism underlying the ZnO micro-javelin formation. Fig. 3 presents the SEM images of samples prepared by irradiating the solutions between 11 min and 22 min and the observations made are summarised in Table 1. The samples prepared at 11 and 12 min showed flaky plate-like structures without any particular orientation or stacking. The 13 min irradiated sample indicated a shaping up to spherical rod-like kind of morphology. On the other hand that prepared with 14 min of irradiation (sample 5 in Table 1) showed the presence of ZnO micro-javelins, and a similar morphology was retained up to 24 min of irradiation.

FTIR spectra of the samples also demonstrate the structural evolution of the ZnO in terms of its shape and purity with the



**Fig. 3** SEM images of the samples obtained after irradiation for (a) 11 min, (b) 13 min, (c) 14 min, (d) 17 min, (e) 19 min and (f) 22 min at a power of 150 W. A typical half-javelin is shown in the box in (d).

**Table 1** Growth of ZnO materials subjected to microwave irradiation between 10 and 24 min

Sample	Irradiation time/min	pH	Dominant morphology	Length of particles	Width of particles/ $\mu\text{m}$	Aspect ratio	Yield (%)
1	10	6.11	—	—	—	—	—
2	11	6.81	Precipitate	—	—	—	5
3	12	6.88	Precipitate	—	—	—	41
4	13	7.04	Precipitate	—	—	—	96
5	14	7.38	Javelins	$16.8 \pm 0.6$	$1.2 \pm 0.1$	14.0	95
6	15	8.62	Javelins	$13.0 \pm 0.5$	$0.9 \pm 0.1$	14.4	95
7	16	8.72	Javelins	$17.9 \pm 0.6$	$1.4 \pm 0.1$	12.8	94
8	17	9.01	Javelins	$14.9 \pm 0.9$	$1.1 \pm 0.1$	13.5	93
9	18	9.05	Javelins	$15.6 \pm 0.4$	$1.1 \pm 0.1$	14.2	92
10	19	9.13	Javelins	$14.6 \pm 0.9$	$1.1 \pm 0.1$	13.3	82
11	20	9.31	Javelins	$14.2 \pm 1.0$	$1.1 \pm 0.1$	12.9	84
12	22	9.56	Javelins	$16.1 \pm 0.7$	$1.1 \pm 0.1$	14.6	82
13	24	9.75	Javelins	$14.0 \pm 0.8$	$1.1 \pm 0.1$	12.7	80

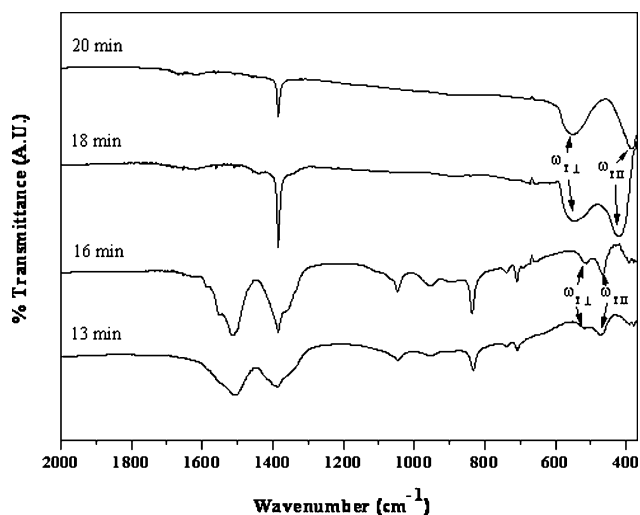


Fig. 4 FTIR spectra of ZnO particles obtained at different MW irradiation times.

time of MW irradiation (Fig. 4). FTIR spectra of samples irradiated for shorter times (those less than 16 min) showed low intensity peaks possibly corresponding to the two transverse optical modes of ZnO wurtzite structure ( $\omega_{T||}$  at 473  $\text{cm}^{-1}$  and 467  $\text{cm}^{-1}$ , and  $\omega_{T\perp}$  at 519  $\text{cm}^{-1}$  and 514  $\text{cm}^{-1}$  for the 13 and 16 min samples)<sup>35</sup> while the longer time irradiated samples showed the presence of well developed spectral bands ( $\omega_{T||}$  at 418  $\text{cm}^{-1}$  and 381  $\text{cm}^{-1}$ , and  $\omega_{T\perp}$  at 547  $\text{cm}^{-1}$  and 551  $\text{cm}^{-1}$  for the 18 and 20 min samples).<sup>36</sup> These observations are consistent with the report by Verges *et al.*, where the evolution of rod shaped particles from spheres was followed using FTIR.<sup>37</sup> The presence of surface-bound species such as adsorbed or complexed species like  $\text{CO}_3^{2-}$ ,  $\text{NO}_3^-$  and surface  $-\text{OH}$  species is identified from their corresponding spectral bands. Bands at 710  $\text{cm}^{-1}$  and 739  $\text{cm}^{-1}$  probably correspond to the  $\nu_4\text{CO}_3^{2-}$  vibrations and the one at 836  $\text{cm}^{-1}$  corresponds to the  $\nu_2\text{CO}_3^{2-}$  vibration.<sup>37,38</sup> All the samples showed a band at 1385  $\text{cm}^{-1}$  indicating the presence of nitrate species complexed with the metal ion.<sup>39</sup> The bands at 1360, 1550 and 1587  $\text{cm}^{-1}$  may be assigned to the symmetric and antisymmetric vibrations of carbonate.<sup>40,41</sup> Further, a broad band centered at 3370  $\text{cm}^{-1}$  ( $\nu_{\text{O-H}}$ ) can be observed for the shorter time irradiated samples (13 and 16 min) indicating the presence of surface adsorbed water molecules (ESI† 3). By contrast, the band corresponding to the 18 min irradiated sample showed only a weak band centered at 3480  $\text{cm}^{-1}$ . Thus these FTIR results confirm that extended ( $\geq 18$  min) microwave irradiation leads to the formation of ZnO particles with elongated morphology and with less adsorbed/complexed species.

### Influence of pH and $\text{Zn}^{2+}$ ion concentration

The pH during the microwave reaction has been monitored (Table 1 and Fig. 5). The solutions obtained on shorter times ( $\leq 10$  min) of MW irradiations, which did not cause ZnO precipitation, showed lower pH values ( $\text{pH} < 6.8$ ) compared to the other samples ( $\text{pH} > 6.8$ ). The solution irradiated for 11–13 min, where precipitates had started to form, gave a  $\text{pH} 6.81$ – $7.04$ . This indicates  $\text{pH} \approx 7$  as the ‘critical solution pH’ for the

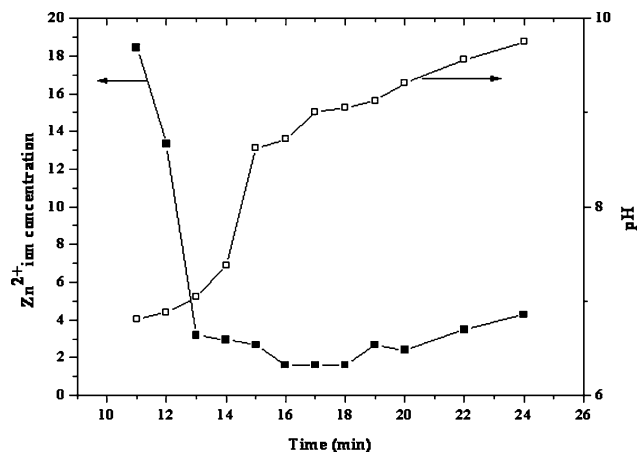


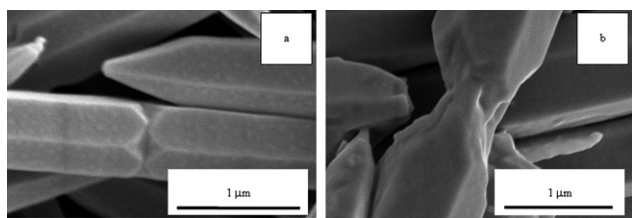
Fig. 5 Plots of pH and  $\text{Zn}^{2+}$  ion concentration vs. MW irradiation time at a power of 150 W.

formation of  $\text{Zn}_5(\text{CO}_3)_2(\text{OH})_6$  (1) particles.<sup>42</sup> Further, the solution pH showed a gradual increase with the extent of MW irradiation and the particles are completely transformed to wurtzite ZnO at a  $\text{pH} > 7.04$ .<sup>42</sup> The pH dependent ZnO precipitation has previously been reported and our results are consistent with those reports.<sup>43–45</sup>

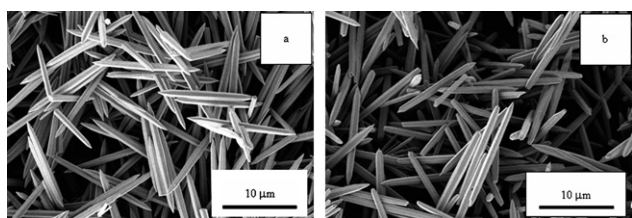
A correlation between microwave irradiation time, pH and  $\text{Zn}^{2+}$  ion concentration was derived by determining the soluble  $\text{Zn}^{2+}$  ion concentration by means of the Zn–EDTA titration of respective microwave irradiated solutions. Fig. 5 presents the profiles showing the evolution of pH and soluble zinc species as a consequence of increasing the MW irradiation time from 11 min to 24 min. A higher soluble zinc species concentration can be observed at pH values less than 7 and as the pH goes up crossing the ‘critical solution pH’ *i.e.*  $\text{pH} \approx 7$ , a sudden decrease in the soluble zinc species concentration occurs. The lowest concentration of soluble zinc species is observed at pH values of  $\sim 8.7$  to 9.05, which can be assumed as the equilibrium concentration, and their concentration in the solution increases somewhat with a further increase in pH. The occurrence of higher concentrations of soluble zinc species in solutions irradiated for longer times could be ascribed to their dissolution (possibly as  $\text{Zn}(\text{OH})_4^{2-}$ ) as a result of an increase in pH beyond that of the point of zero charge (PZC) of ZnO (*ca.* 9).<sup>18</sup>

It may be noted that the samples especially those prepared with  $> 17$  min of MW irradiation showed the presence of some half-javelins. An etching from the middle portion of the javelins can also be observed in such samples. This could be attributed to the corresponding pH of the microwave treated solution (Table 1). When the particles are cooled in a solution with a pH near or above the PZC of ZnO particles, etching of  $\text{Zn}^{2+}$  may occur. Such a preferential etching is found to be happening from the axially middle portion of the ZnO micro-javelins (Fig. 6), which indicates it is a more susceptible region compared to other regions.

In order to confirm the influence of pH on the solubilisation of the zinc species/etching of the ZnO crystals, a number of further experiments have been carried out. Firstly the pH of the zinc nitrate–urea mixture was adjusted to 10 using  $\text{NH}_4\text{OH}$  solution and then irradiated for 18 min. In a second experiment, the reactant mixture was first irradiated for 14 min, allowed to cool



**Fig. 6** SEM images showing the etching from the middle portion of the javelins (a) after 20 min reaction and (b) after 21 min at a power of 150 W.



**Fig. 7** (a) ZnO produced after immediate pressure release and (b) pressure released after cooling to room temperature. Samples prepared by reaction at 150 W for 20 min.

down to room temperature, the pH adjusted to 10 with  $\text{NH}_4\text{OH}$  and then the MW irradiation was carried out for a further 18 min. Both the experiments produced micro-javelins with etched middle portion and a rough surface texture (ESI† 4). A corresponding increase in zinc species concentration was also observed in the solution substantiating the role of solution pH.

#### Effect of pressure release

In order to understand the effect of pressure in the Teflon reaction vessel on the size and/or shape of the product, two types of experiments were initially conducted. In our preliminary studies<sup>30</sup> the pressure was released immediately after the MW irradiation whereas in the current study the vessel was allowed to cool to room temperature and only at that stage was the pressure reduced by opening the reactor vessel valve. Both procedures gave (1000) oriented crystalline ZnO micro-javelins as the product, but the second procedure produced slightly shorter ZnO particles ( $\sim 14.2 \mu\text{m}$ ) compared to the first procedure ( $\sim 15.7 \mu\text{m}$ ) (Fig. 7). (In both cases the diameter is  $1.1 \mu\text{m}$ .) Such decrease might be attributable to the solubilisation of  $\text{Zn}(\text{OH})_4^{2-}$  species as a result of the prolonged aging (20 h) in the basic solution (pH > 9). It should be noted that factors such as the free volume in the vessel and the oxygen partial pressure in the vessel may also play important roles.

#### Effect of counterions

Reactions with different zinc precursors such as zinc acetate, zinc chloride and zinc sulfate have also been carried out to investigate and compare the effect of counterions on the micro-javelin formation. Zinc acetate gave multidisperse ZnO half-javelins of  $\sim 4 \mu\text{m}$  length (Fig. 8a). Zinc chloride produced ZnO micro-javelins of  $18 \pm 2 \mu\text{m}$  length (Fig. 8c). Zinc sulfate on the other hand produced mostly agglomerated particles with some half-javelins (Fig. 8e). Corresponding XRD patterns indicated the

formation of crystalline wurtzite ZnO by both acetate and chloride routes (Fig. 8b and d), while the sulfate route (Fig. 8f) gave a mixture of zinc carbonate and ZnO peaks but with lower intensity. The comparatively higher complexing ability of acetate may be the reason behind the poor formation of ZnO micro-javelins from  $\text{Zn}(\text{OAc})_2$ ,<sup>46</sup> compared to the counterions such as  $\text{Cl}^-$  and  $\text{NO}_3^{2-}$ , which have less significant influence on the nucleation and growth of the metal carbonate and oxide particles.

#### Effect of urea/ammonium carbamate

In order to investigate the role of urea hydrolysis products, experiments were conducted with 100 mM ammonium carbamate (instead of urea) and 10 mM zinc nitrate solution for 20 min. Interestingly, a wire-like particle morphology was observed in the SEM (Fig. 9). The wires were found to be composed of  $\text{Zn}_4\text{CO}_3(\text{OH})_6 \cdot \text{H}_2\text{O}$  rather than ZnO particles. On calcination to  $500^\circ\text{C}$  the particles were converted to ZnO with capsule like morphology (ESI† 5). This experiment suggests that carbamate ions are less likely to play a major role in the formation of the micro-javelins.

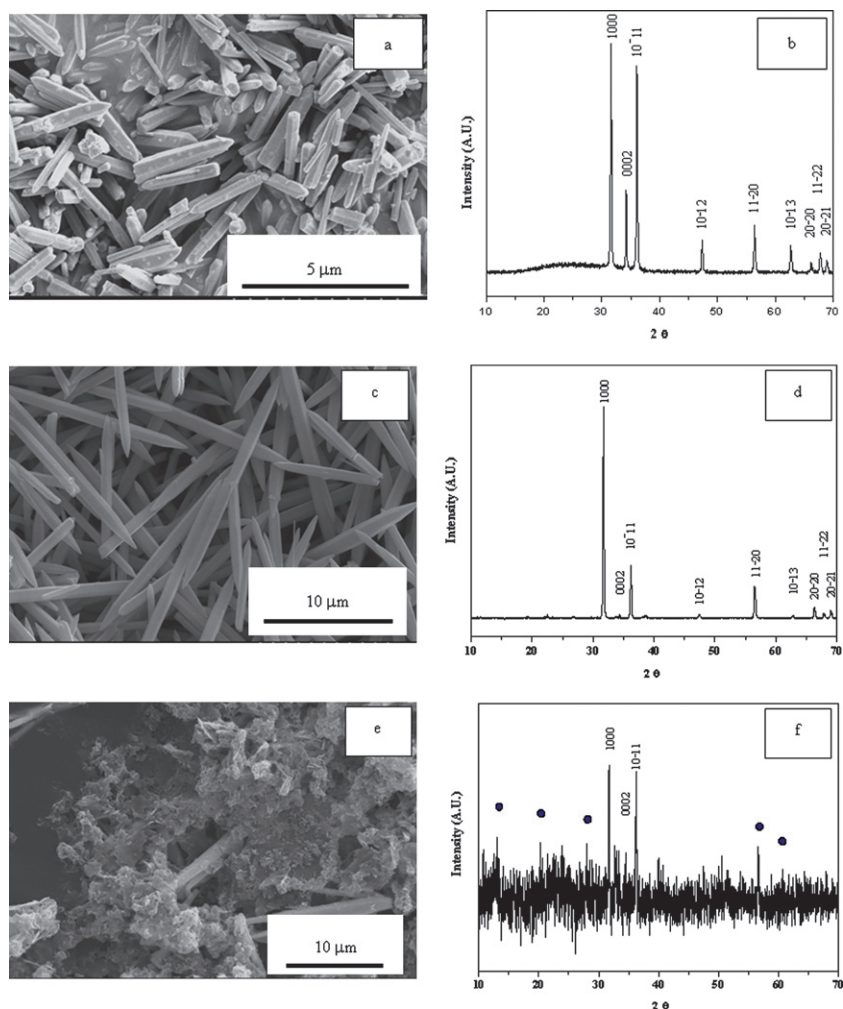
#### Photoluminescence of ZnO micro-javelins

Fig. 10 presents the room temperature PL spectra of ZnO particles obtained after 13, 18 and 24 min of MW irradiation. All the spectra show a near band edge emission at  $\sim 3.2 \text{ eV}$ .<sup>47–50</sup> Further, an enhancement in intensity of the 390 nm band can be observed with increasing MW irradiation time. The defect luminescence band in the green region, which is the case for most of the reported ZnO particles,<sup>51</sup> could not, however, be observed in the present samples. This method could therefore be assumed as giving reasonably defect-free ZnO particles that could have wide technological implications in LEDs, lasers and transparent transistors.<sup>52</sup>

#### Discussion

Highly crystalline, monodisperse ZnO micro-javelins with dominant (1000) diffraction have been obtained by a one-step MW-assisted synthesis procedure. An investigation on the effect of pH,  $\text{Zn}^{2+}$  ion concentration and time of MW irradiation reveals the structural evolution of javelins. XRD of the 13 min irradiated sample showed the diffractions corresponding to a mixture of small fractions of ZnO and  $\text{Zn}_5(\text{CO}_3)_2(\text{OH})_6$  (I)<sup>42</sup> while all the other samples (irradiated for longer time) produced wurtzite ZnO. The intermediate zinc carbonate compound (I) and ZnO nuclei are precipitated as a consequence of the MW induced supersaturation of soluble zinc species and the release of species such as  $\text{HCO}_3^-$  and  $\text{NH}_3$  into the solution. On continuing the irradiation further, the intermediate carbonate compound redissolves/dissociates and the growth units are formed in turn, which then add onto the existing ZnO nuclei to form the javelins. Urea being a comparatively neutral hydrolysis medium also favours the elongated prismatic growth of ZnO particles probably due to its comparatively weak complexing ability to the  $\text{Zn}^{2+}$  species. In comparison, alkaline NaOH and other strong complexing agents are reported to suppress the axial growth of ZnO





**Fig. 8** (a) FESEM and (b) XRD of sample prepared by the zinc acetate and urea at 150 W for 20 min. (c) FESEM and (d) XRD of sample prepared by the zinc chloride and urea at 150 W for 20 min. (e) FESEM and (f) XRD of sample prepared by the zinc sulfate and urea at 150 W for 20 min. ●—  $\text{Zn}_5(\text{CO}_3)_2(\text{OH})_6$  (1).

particles by strongly complexing with the surface  $\text{Zn}^{2+}$  ions in the (0001) face.<sup>53</sup>

The SEM experiments clearly demonstrate the change of shape as the experiment proceeds. Thus the 11 and 12 min irradiated samples show amorphous precipitates with flaky plate-like morphology, whereas in the 13 min sample rod-like particles are present.  $\text{ZnO}$  micro-javelins were already present in the 14 min irradiated sample and this structure was found to be retained on further irradiation up to 24 min. Further, it may be noted that samples prepared by irradiating the solution for >17 min showed an interesting etching behaviour through the middle portion of the micro-javelins (Fig. 6). This indicates the possibility of the presence of a zinc-rich region towards the middle of the micro-javelins.

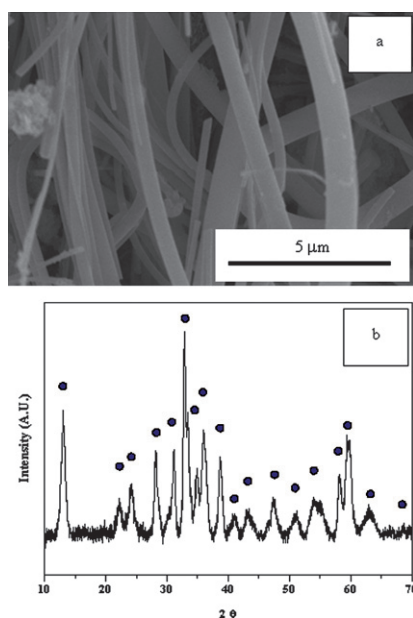
### Growth mechanism

It is essential to know the structural habit of  $\text{ZnO}$  to understand the growth mechanism. The structure of  $\text{ZnO}$  can be described as consisting of hexagonal close packing of oxygen and zinc atoms in space group  $P6_3mc$  with oxygen atoms bonded to zinc atoms

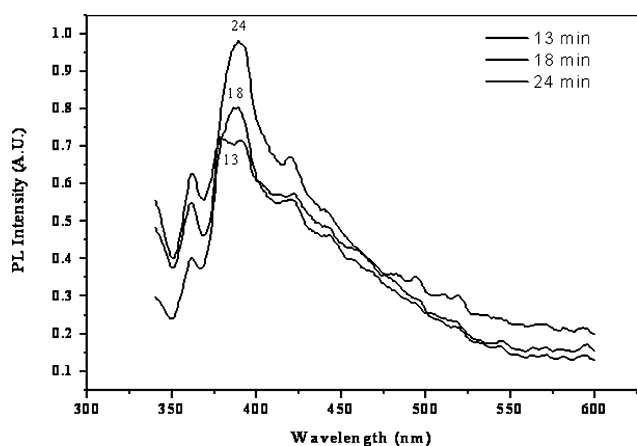
occupied in tetrahedral sites.<sup>54</sup> The zinc and oxygen atoms are alternately arranged along the  $c$ -axis of the hexagonal phase. The structure has no centre of inversion, and this anisotropy provides an inherent polarity to the crystal along the  $c$ -axis.

The (0001) plane (terminated with  $\text{Zn}^{2+}$  ions) of  $\text{ZnO}$  has the maximum surface energy, while the (000-1) plane (terminated with  $\text{O}^{2-}$  ions) has the minimum surface energy. The faces with the moderate growth velocities are  $V_{(-101-1)} > V_{(-1010)} > V_{(-1011)}$ .<sup>54</sup> Schematic presented in Fig. 11 shows various faces of a typical javelin-like  $\text{ZnO}$  particle.

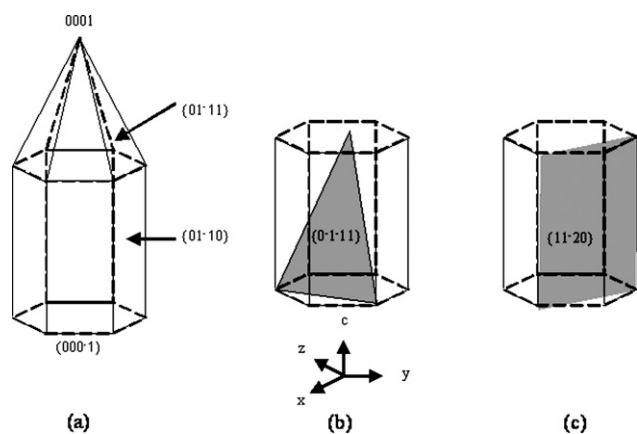
The formation of bridged  $\text{ZnO}$  structures has been previously observed.<sup>33,34,55</sup> Oliveira *et al.*<sup>33</sup> and Xie *et al.*<sup>34</sup> proposed a germination of the second half from the base of a preformed half-ellipsoid as the mechanism of ellipsoid formation. Further, it has been shown in several examples that the decomposition of urea in aqueous solutions at elevated temperatures can be utilised to produce colloidal dispersions consisting of uniform particles.<sup>56</sup> The influence of zinc counterions on the shape evolution has also been reported.<sup>18</sup> Literature on the pH dependence of  $\text{ZnO}$  and  $\text{Zn}(\text{OH})_2$  formation is also available.<sup>43</sup> However, to the best of our knowledge, the mechanism underlying the formation of  $\text{ZnO}$



**Fig. 9** (a) SEM and (b) XRD of sample prepared from zinc nitrate and ammonium carbamate. ●—Zn<sub>5</sub>(CO<sub>3</sub>)<sub>2</sub>(OH)<sub>6</sub>.



**Fig. 10**

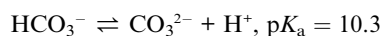
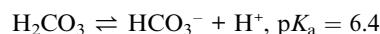
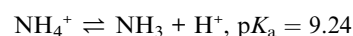


**Fig. 11** Schematic showing various faces of the hexagonal ZnO particle: (a) half-javelin, (b) (0-1-11) face in a unit cell and (c) (11-20) face in a unit cell.

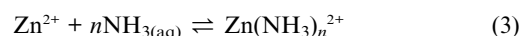
micro-javelins under the restricted pH condition has not been reported previously.

In much of the literature presenting hydrothermal synthesis of ZnO, Zn(OH)<sub>4</sub><sup>2-</sup> was reported as the growth unit for the formation of wurtzite ZnO.<sup>57,58</sup> However, in the present rapid microwave process in the presence of urea, where the pH of the solution is found to be restricted to <10, Zn<sup>2+</sup>, Zn(OH)<sup>+</sup> and Zn(OH)<sub>2</sub> could be the possible zinc species<sup>43,59</sup> with highest supersaturation. Hence the growth process may be different. Indeed, the micro-javelins are formed rapidly under weakly basic conditions (pH 7–7.5) by the present microwave process.

In the present case, it is likely that a number of molecular species influence this reaction. It is known that upon heating urea breaks down to form initially carbamate and ammonia. Subsequent reaction leads to carbonate/bicarbonate formation.<sup>60</sup> As a result of the formation of these species the pH rises as is indeed observed during our experiments.

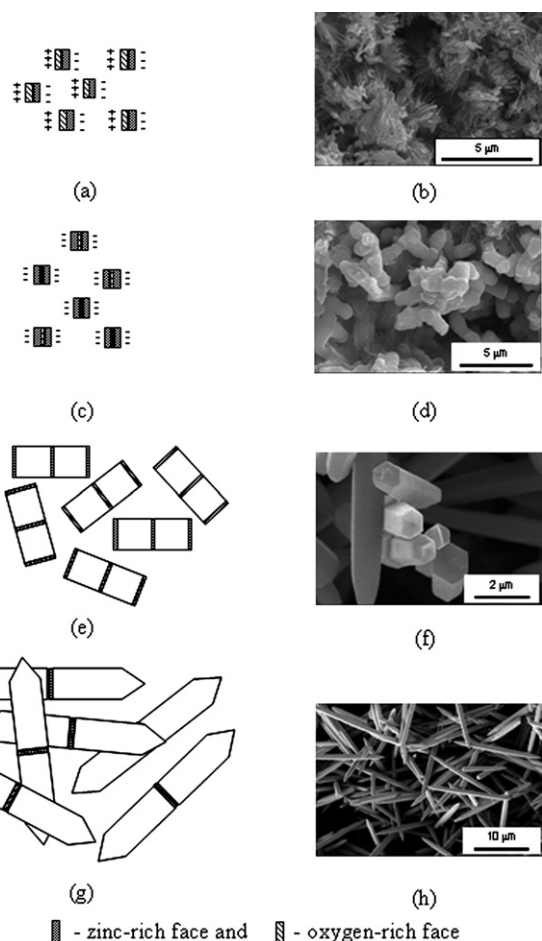


Additionally, under our reaction conditions the ammonia may complex with zinc ions



In our process the solid forming initially (*e.g.* at 11 min) contains a large amount of Zn<sub>5</sub>(CO<sub>3</sub>)<sub>2</sub>(OH)<sub>6</sub> (**1**). However, at times >14 min ZnO is the only solid species identified. It may be noted that **1** decomposes at temperature >150 °C. This may occur in our case or alternatively **1** may redissolve. We favour the redissolution of **1** as this should happen with the increase in NH<sub>3</sub> concentration. The formation of ZnO is consistent with the increase in pH at longer time.

Fig. 12 presents a schematic of the events happening in the growth of ZnO micro-javelins. In the case of the 11–12 min samples the growth of nuclei is inhibited possibly by its coordination with the (NH<sub>2</sub>)<sub>2</sub>COO<sup>-</sup> species produced as a result of the decomposition of urea. Further irradiation induces the decomposition of (NH<sub>2</sub>)<sub>2</sub>COO<sup>-</sup> species into NH<sub>4</sub><sup>+</sup> and CO<sub>3</sub><sup>2-</sup> species, which results in the formation of thermodynamically stable Zn<sub>5</sub>(CO<sub>3</sub>)<sub>2</sub>(OH)<sub>6</sub> particles (ESI† 2).<sup>61</sup> The hydrozincite (**1**)



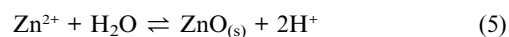
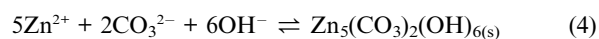
**Fig. 12** Schematic showing the sequential processes occurring in the ZnO javelin formation. 11–12 min irradiated samples (particles with a flaky shape) are saturated with urea-complexed  $Zn^{2+}$  species [nuclei having urea coordination ( $NH_2COO^-$ ) at the Zn-rich face (a) and (b)]. 13 min of irradiation induces an initial growth through the oxygen-rich face by the deposition of  $Zn(OH)^+$  species [linking of particles (c) and (d)]. On further irradiation, the deposition of  $Zn(OH)_2$  occurs by dehydration with surface Zn–OH [(e) and (f)] and javelins are formed subsequently [ $\geq 14$  min (g) and (h)].

particles, however, are found to be unstable under the presence of ammonia, which is being generated subsequently. The dissolution of hydrozincite particles (1) occurs as a result and the formed  $Zn(OH)^+$  species subsequently add onto the oxygen-rich face of the nuclei (linking of particles).<sup>19,43,62</sup> This results in the formation of a spherical- to rod-like kind of morphology.

According to the phase stability diagrams for the ZnO–H<sub>2</sub>O and Zn(OH)<sub>2</sub>–H<sub>2</sub>O systems at 25 °C and thermodynamic data (ESI† 6) presented in ref. 43, the soluble species such as  $Zn^{2+}$  and  $ZnOH^+$  are stable below pH  $\approx 7$  at a  $Zn^{2+}$  concentration of 5 mM. These diagrams also indicate that the solubility of ZnO is almost the same as that of Zn(OH)<sub>2</sub>. Thus, both Zn(OH)<sub>2</sub> and ZnO should be thermodynamically stable over a pH range of 8 to 12 at a 5 mM  $Zn^{2+}$  concentration. The wurtzite ZnO particles are, however, observed, in the present case, at the pH range 7.04–9.75 and are consistent with the phase stability diagram. This selective crystallisation and growth of wurtzite ZnO particles could also

be attributed to the presence of existing ZnO nuclei, which act as heterogeneous nucleation sites.

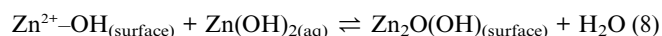
Castellano and Matijevic previously reported the formation of rod shaped basic zinc carbonate (hydrozincite) particles from aqueous solutions of zinc nitrate and urea at elevated temperatures (90 °C) and its subsequent transformation to ZnO rods on calcination.<sup>63</sup> Our experiments resulted in the formation of a mixture of zinc carbonate and ZnO structures at the initial stages of MW irradiation. With increasing irradiation time, the ZnO particles are grown through an intermediate amorphous Zn(OH)<sub>2</sub> pathway. Such a reaction sequence also corroborates with a previous observation by McBride *et al.*, where the formation of intermediate Zn(OH)<sub>2</sub> has been suggested as a prerequisite for the controlled growth of ZnO needles.<sup>16</sup> However, X-ray diffractions corresponding to Zn(OH)<sub>2</sub> could not be observed in any of our low irradiation time samples. In light of all these factors, it is proposed that the ZnO javelin growth occurs through a different pathway under the conditions of the effective and rapid homogeneous heating of the microwave process. Based on our experimental results and the previous reports<sup>25,43,57,64,65</sup> the growth process can be explained as follows. The high supersaturation of  $Zn^{2+}$  species at the initial stage (under the neutral pH condition) of the MW irradiation of zinc nitrate and excess urea (10 times) results in the precipitation of  $Zn_5(CO_3)_2(OH)_6$  and ZnO nuclei according to the reaction:



Further irradiation results in the building up of NH<sub>3</sub> in the solution. Under such weakly alkaline condition,  $Zn^{2+}$  cations form ammine complexes of  $Zn(NH_3)_n^{2+}$  as in eqn (3). The ammine complex formation shifts the equilibrium in reaction (4) to the left. This in turn results in the supersaturation of species such as  $Zn(OH)^+$  and Zn(OH)<sub>2</sub> in the solution according to the reaction:



All these sequential reactions are consistent with the pK<sub>a</sub> values. This sequential formation of  $Zn(OH)^+$  under the weakly alkaline condition appears to be the key step that initiates the formation of javelin-like particles in the present system. The addition of  $Zn(OH)^+$  species onto the oxygen-rich (000–1) face of the existing ZnO nuclei results in the formation of a structure which is rich in  $Zn^{2+}$  species on both the ends of the particles (Fig. 12c). Subsequently, javelins are formed by the dehydration reaction between the surface hydroxide species (either end) and the OH<sup>−</sup> species on the zinc complexes according to reaction:



The structural fit between the phases favours the heterogeneous nucleation of the supersaturated Zn(OH)<sub>2</sub> species to the ZnO nuclei rather than the bulk homogeneous nucleation and growth of ZnO nuclei.<sup>43,65</sup> Further, it can be observed that the micro-javelins form at a moderately alkaline pH (sample 5 in

Table 1). Hence, the supersaturation of species such as  $\text{Zn(OH)}^+$  and  $\text{Zn(OH)}_2$  should be the defining factors here for micro-javelins growth. The ammine complexation does not seem to have a strong influence on the micro-javelin formation, due to the lower  $\text{NH}_4^+$  to  $\text{Zn}^{2+}$  ratio in the present case, as previously reported by Yamabi and Imai.<sup>43</sup> However, it plays a key role in the dissolution of the initially formed hydrozincite particles. Further, the pH dependent bicarbonate–carbonate equilibria (ESI† 7) indicate bicarbonate as the predominant species in the pH range 8–10, which is consistent with the FTIR results.

The present microwave-assisted process resulted in the formation of highly crystalline monodispersed micro-javelins with significantly larger sizes ( $\sim 15 \pm 2 \mu\text{m}$  length vs.  $\sim 1 \mu\text{m}$  width) compared to the previous reports.<sup>66</sup> This could possibly be attributed to the rapid homogeneous heating of the MW process, which leads to a greater extent of supersaturation of ZnO and thus results in a small number of ZnO nuclei. The growth onto these limited numbers of nuclei therefore results in larger particles. Indeed, the unusual XRD pattern shown by the particles is attributed to such a high aspect ratio, which when placed in a glass plate (substrate used for XRD) tends to align themselves with their *c*-axis parallel to the substrate. Such an arrangement enhances the probability for most of their (1000) planes to come against the X-ray beam, compared to any other planes, hence giving a high intensity peak.

The optical studies of the micro-javelins showed only a weak PL emission. However, the striking observation was the absence of green light defect luminescence. This can be attributed to relatively low amounts of surface defects for the longer time irradiated samples. The present method is therefore found to be effective in producing reasonably defect-free ZnO particles.

## Conclusions

In summary, a straightforward one-step microwave reaction between zinc nitrate and urea in aqueous solutions produced highly crystalline wurtzite ZnO micro-javelins with a predominant (1000) X-ray diffraction. The javelins have, however, been found to grow in the (0001) direction and the appearance of the extremely dominant (1000) X-ray diffraction peak is attributed to their high aspect ratio. The growth is proposed to involve the MW induced precipitation of ZnO nuclei and metastable zinc carbonate structures along with the release of species such as  $\text{NH}_2\text{COO}^-$ ,  $\text{HCO}_3^-$  and ammonia into the solution. Upon continued irradiation, the dissolution of the intermediate carbonate structures results in the formation of species such as  $\text{Zn(OH)}^+$  and  $\text{Zn(OH)}_2$  and the addition of  $\text{Zn(OH)}^+$  onto the oxygen-rich face of the existing ZnO nuclei occurs to form rod-like kind of particles having an outermost  $\text{Zn}^{2+}$  layer at both its ends. Such particles then grow to form the micro-javelins by a heterogeneous nucleation-assisted growth between them and the  $\text{Zn(OH)}_2$  species. The consumption of  $\text{NH}_2\text{COO}^-$  species upon extensive irradiation results in the elimination of the growth barrier that in turn dictates the tapering of both the ends of the particles to form pointed tips, hence micro-javelins.

The microwave induced supersaturation of  $\text{Zn}^{2+}$  and  $\text{Zn(OH)}^+$  species under the moderately basic pH condition, and the resultant initial growth through the oxygen terminated (000–1) facet, have been identified as the key steps responsible for the

formation of micro-javelins. An enhanced UV emission was shown by such perfectly grown ZnO javelins with reasonably low green luminescence. Furthermore, the method does not involve any complex additives, surfactants or polymeric agents. Well-crystallised products are formed in a single-step process without the need for special separation procedures. This straightforward method can therefore be used for the preparation of reasonably defect-free ZnO micro-javelins in a cost-effective way.

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