

2001

Preparation of Magnetic Nanoparticles and their Assemblies Using a New Fe(II) Alkoxide Precursor

Suresh Pillai

Technological University Dublin, suresh.pillai@tudublin.ie

Grainne Biddlecombe


Yurii Gun'ko

John Kelly

J. Michael Coey

See next page for additional authors

Follow this and additional works at: <https://arrow.tudublin.ie/cenresart>

 Part of the [Ceramic Materials Commons](#), [Nanoscience and Nanotechnology Commons](#), and the [Other Materials Science and Engineering Commons](#)

Recommended Citation

Biddlecombe, G. (2003) Preparation of magnetic nanoparticles and their assemblies using a new Fe(II) alkoxide precursor. *Journal of Materials Chemistry*, 11, 2001, 2937-2939. doi:10.1039/B107760K

This Article is brought to you for free and open access by the Crest: Centre for Research in Engineering Surface Technology at ARROW@TU Dublin. It has been accepted for inclusion in Articles by an authorized administrator of ARROW@TU Dublin. For more information, please contact yvonne.desmond@tudublin.ie, arrow.admin@tudublin.ie, brian.widdis@tudublin.ie.



This work is licensed under a [Creative Commons Attribution-Noncommercial-Share Alike 3.0 License](#)

Authors

Suresh Pillai, Grainne Biddlecombe, Yurii Gun'ko, John Kelly, J. Michael Coey, and Alexis Douvalis

Preparation of magnetic nanoparticles and their assemblies using a new Fe(II) alkoxide precursor

Grainne B. Biddlecombe,^a Yurii K. Gun'ko,^{*a} John M. Kelly,^a Suresh C. Pillai,^a
 J. Michael D. Coey,^b Munuswamy Venkatesan^b and Alexis P. Douvalis^b

^aThe Department of Chemistry, Trinity College, University of Dublin, Dublin 2, Ireland.

E-mail: igounko@tcd.ie

^bThe Department of Physics, Trinity College, University of Dublin, Dublin 2, Ireland

Received 29th August 2001, Accepted 3rd October 2001

First published as an Advance Article on the web 15th October 2001

Hydrolysis of the new metallorganic precursor $\text{Fe}(\text{OBU})_2(\text{THF})_2$ followed by ultrasound and thermal treatment yielded either nanoparticles of $\gamma\text{-Fe}_2\text{O}_3$ or Fe_3O_4 , depending on the process conditions. The samples have been characterised by X-ray powder diffraction, TEM, SEM, Mössbauer spectroscopy and magnetization measurements. The maghemite nanoparticles (9 ± 2 nm) are superparamagnetic and form unique needle-like assemblies of nanoparticle arrays. The nanoparticles of Fe_3O_4 (19 ± 2 nm) form plate-like aggregates about 10 μm thick.

Magnetic nanoparticles, which approach the size of a single magnetic domain, have applications in magnetic storage.^{1,2} In order to get higher storage densities future development in the area of magnetic storage media will rely on the ability to develop stable materials with particles a few nanometers in size.³ Superparamagnetic single domain nanoparticles also have very important biomedical applications (*e.g.* magnetic cell sorting, magnetocytolysis, drug-targeting experiments, gene transfection studies, magnetic fluid hyperthermia, *etc.*).⁴⁻⁶

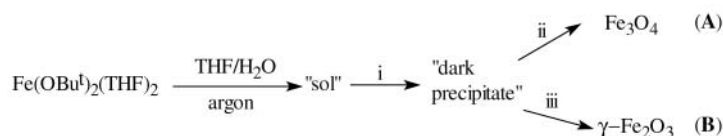
The development of reliable techniques for preparing arrays and assemblies of magnetic nanoparticles is of high scientific and technological importance. Despite the wide range of physical and chemical routes for the preparation of nanoparticles which have been developed over many years,⁷ there is a great demand for precursors and for an effective technique to produce nanoparticles of magnetite (Fe_3O_4) and maghemite ($\gamma\text{-Fe}_2\text{O}_3$). Several new chemical routes for preparation of different magnetic nanoparticles and their arrays have been developed.⁸ Using single-source precursors is one of the most attractive methods for the preparation of magnetic nanoparticles. A particularly interesting non-hydrolytic single-precursor approach for the preparation of metal oxides, including $\gamma\text{-Fe}_2\text{O}_3$ nanoparticles, from the corresponding metal cupferronates, was proposed by Alivisatos and coworkers.⁹ Recently Gedanken and coworkers have also shown that Fe_3O_4 nanoparticles can be effectively prepared using a sonochemical method with commercially available metallorganic precursors such as $\text{Fe}(\text{CO})_5$ and Fe(II) acetate.¹⁰

In this communication we report the preparation of

nanoparticles of magnetite and maghemite ($\gamma\text{-Fe}_2\text{O}_3$) and their assemblies using a new metallorganic iron(II) alkoxide precursor $\text{Fe}(\text{OBU})_2(\text{THF})_2$ by modified sol-gel and ultrasonic techniques.

The precursor, $\text{Fe}(\text{OBU})_2(\text{THF})_2$ was prepared from FeBr_2 and NaOBU^1 in THF using standard argon-vacuum and Schlenk techniques.† A solution of $\text{Fe}(\text{OBU})_2(\text{THF})_2$ (1.1 g, 2 mmol) in THF (60 cm^3) was carefully hydrolysed with doubly distilled non-degassed water (30 cm^3) under argon (Scheme 1). The resultant mixture was then ultrasonically (30 kHz, 130 W) irradiated under argon at ambient temperature for 1 h. The dark precipitate was filtered off and washed several times with distilled water. The product was then divided into two equal parts. One part of the product was dried *in vacuo* for 3 h at room-temperature giving a black powder (A), while the other part was dried in air at 200 °C for 30 min giving a dark brown powder (B). Both products were characterised by X-ray powder diffraction (XRD). Reflections of the XRD patterns (2θ from 5 to 70°) for both samples showed d spacing values and relative intensities of the peaks coincident exactly with JCPDS data of magnetite (Fe_3O_4) for A and maghemite ($\gamma\text{-Fe}_2\text{O}_3$) for B with some line-broadening. The particle sizes were calculated using the Debye-Scherrer formula¹¹ (crystalline diameter = $0.9\lambda/B \cos\theta$, where λ = X-ray wavelength, θ = Bragg angle, B = line broadening) giving a particle diameter of 19 ± 2 nm for Fe_3O_4 and 9 ± 2 nm for $\gamma\text{-Fe}_2\text{O}_3$ samples.

Fig. 1 presents the Mössbauer spectra of the two samples taken at room temperature. For the black powder (A) the spectrum shows combined magnetic and quadrupole splittings. A least square fit to the experimental data indicates that the lines are characteristic of non-stoichiometric magnetite, with a small line broadening due to a distribution of hyperfine magnetic fields caused by the different iron ion environments in the structure, and a small contribution of paramagnetic Fe^{3+} . For the brown powder (B) the room-temperature spectrum shows only quadrupole splitting. A fitting to these data suggests that the quadrupole doublet can be assigned to Fe^{3+} ions in superparamagnetic $\gamma\text{-Fe}_2\text{O}_3$ particles as indicated by the absence of any magnetic splitting.



Scheme 1 Schematic illustration of the preparation of magnetic nanoparticles. Conditions: i, ultrasound, 1 h; ii, vacuum, 3 h, *ca.* 20 °C; iii, heat, *ca.* 200 °C, in air, 30 min.

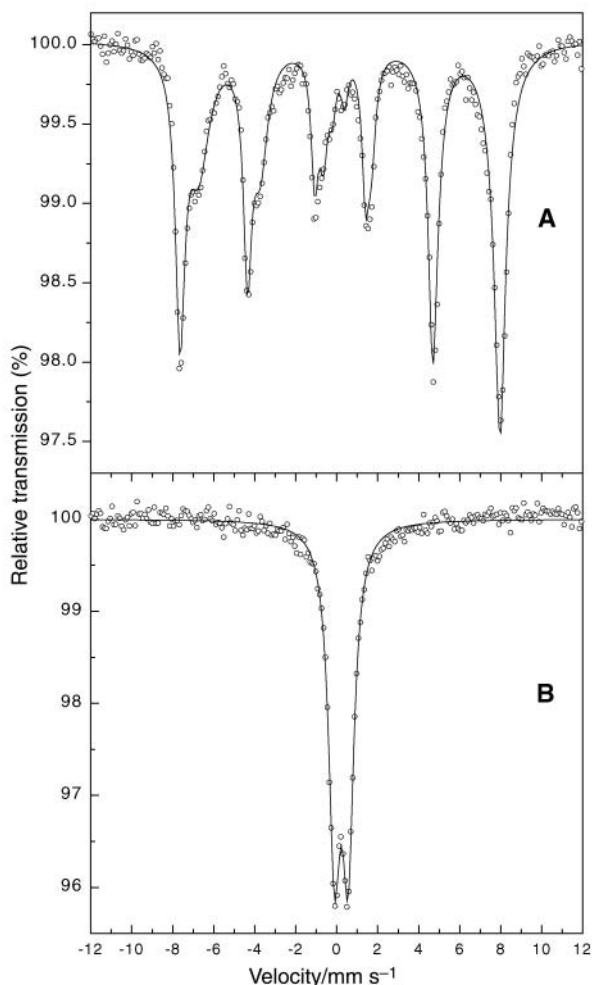


Fig. 1 Room-temperature Mössbauer spectra of Fe_3O_4 (A) and $\gamma\text{-Fe}_2\text{O}_3$ (B) nanoparticles.

Field emission scanning electron microscopy (FESEM) images of the powders demonstrated that the Fe_3O_4 sample (Fig. 2) consists of broken plates, which are about $100\ \mu\text{m}$ in

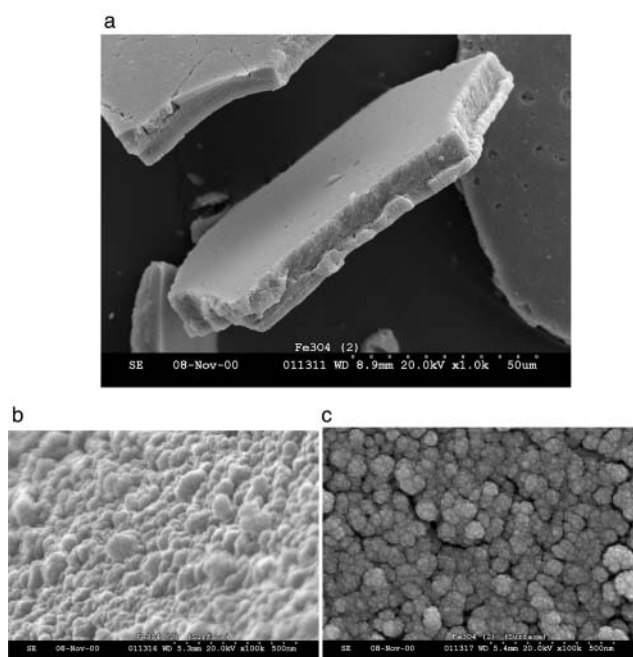


Fig. 2 FESEM microphotographs at $1000\times$ (a) and $100\,000\times$ (b, broken edge; c, plane surface) magnification of the Fe_3O_4 sample (A).

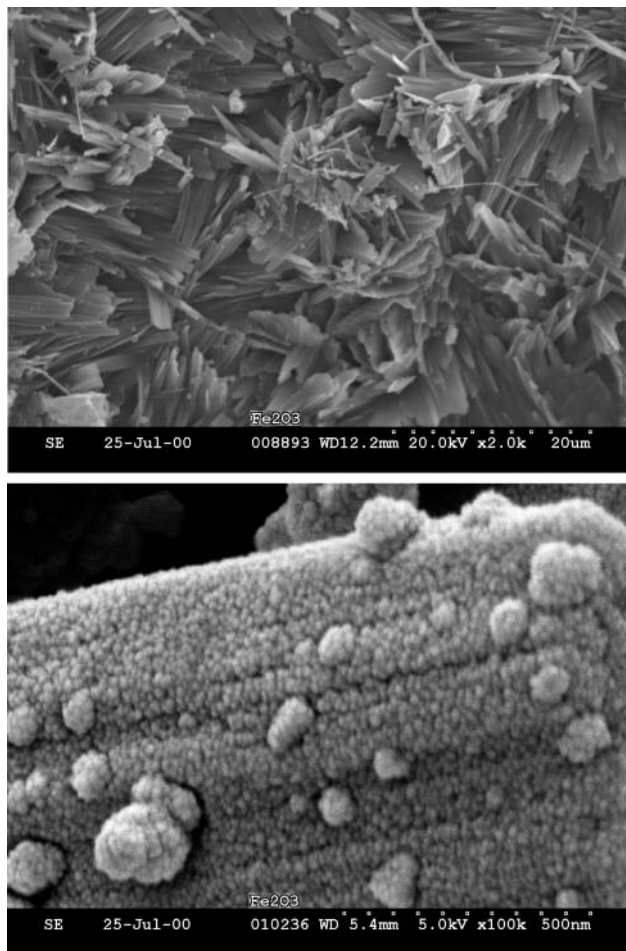


Fig. 3 FESEM microphotographs at $2000\times$ (top) and $100\,000\times$ (bottom) magnification of the $\gamma\text{-Fe}_2\text{O}_3$ needles (B).

diameter and about $10\ \mu\text{m}$ thick. Magnification of the broken edge and plate surface up to $150\,000\times$ shows that each plate is an assembly of nano-sized particles. The estimated size of individual aggregates from FESEM is between $10\text{--}100\ \text{nm}$. By contrast the FESEM images of $\gamma\text{-Fe}_2\text{O}_3$ at $2000\times$ and $100\,000\times$ magnification (Fig. 3) show a straw- or needle-like sample morphology with needle lengths of *ca.* $10\ \mu\text{m}$. The image at $100\,000\times$ indicates that each straw or needle consists of three-dimensional self-assemblies of parallel arrays of nanoparticles. The average size of the each nanoparticle is *ca.* $10\ \text{nm}$.

Samples for transmission electron microscopy (TEM) were prepared by deposition of a drop of very dilute colloidal solutions of Fe_3O_4 or $\gamma\text{-Fe}_2\text{O}_3$ in ethanol onto formvar coated 400 mesh copper grids. TEM images of the magnetite sample show a dispersion of nanoparticle aggregates with a size distribution between 10 and $60\ \text{nm}$. According to TEM images the nanoparticles of the maghemite sample have a quite narrow size distribution and an average diameter of $7\text{--}10\ \text{nm}$.

Fig. 4 shows the room-temperature magnetization of the Fe_3O_4 sample. The room-temperature spontaneous magnetization (σ) is $65\ \text{J T}^{-1}\ \text{kg}^{-1}$ in a field of $1\ \text{T}$ which is considerably less than the usual spontaneous saturation magnetisation (*ca.* $90\ \text{J T}^{-1}\ \text{kg}^{-1}$) of bulk Fe_3O_4 .¹² The value of the coercivity is around $10\ \text{mT}$. A decreasing size of particles leads to a decrease in total magnetization due to the increased dispersion in the exchange integral.¹³ On the other hand, the magnetization curve of the maghemite sample is typical of superparamagnetic behaviour showing a steep increase in magnetisation with increasing field. The magnetisation does not saturate even at $1\ \text{T}$ magnetic field. Using the expression for the mass susceptibility of a paramagnet ($\chi = Nm^2/3k_B$, where N is the

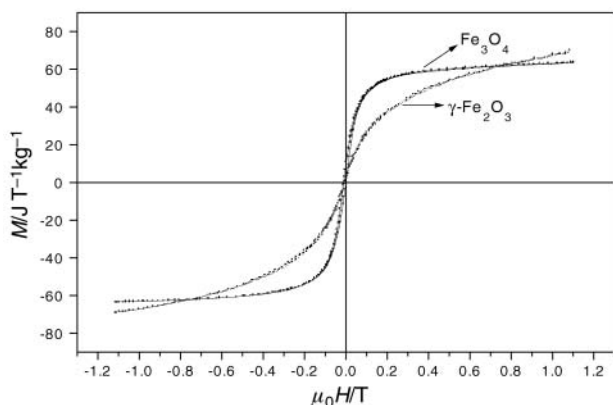


Fig. 4 Room-temperature magnetization curves of Fe_3O_4 and $\gamma\text{-Fe}_2\text{O}_3$ samples.

number of particles per kg and m is the particle moment (J T^{-1}), from the initial slope a particle diameter of 7 ± 1 nm is deduced. This size is quite consistent with XRD, FESEM and TEM results. The lack of saturation of magnetisation and the reduction in the room-temperature magnetisation may be attributed to the smaller particle size.

The mechanism of formation of magnetic nanoparticles obviously involves an oxidation of Fe(II) into Fe(III) by oxygen from the water during the hydrolysis and further ultrasonic treatment. Ultrasonically generated radical and peroxide species from water or THF also might contribute to the oxidation of Fe(II) , as has been recently reported by Gedanken and coworkers.^{9,10} We suggest that the formation of $\gamma\text{-Fe}_2\text{O}_3$ nanoparticles proceeds *via* precursor materials containing $-\text{Fe}(\text{OH})-\text{Fe}-$ and $-\text{Fe}-\text{O}-\text{Fe}-$ polymeric chains and that these then can be ultrasonically and thermally broken giving individual nano-sized aggregates. Plausibly even after the treatment, aggregates still continue to retain a chain-like orientation forming needle-like assemblies of nanoparticles as observed by FESEM.

In summary, the results illustrate that nanoparticles of magnetite (Fe_3O_4) and maghemite ($\gamma\text{-Fe}_2\text{O}_3$) can be conveniently produced from an Fe(II) alkoxide derivative. The maghemite nanoparticles so prepared are superparamagnetic. Results of FESEM have shown that nanoparticles of Fe_3O_4 form plate-like aggregates of $100 \mu\text{m}$ diameter and about $10 \mu\text{m}$ thickness; $\gamma\text{-Fe}_2\text{O}_3$ consists of needle-like assemblies of nanoparticle arrays. This work presents new opportunities for the

preparation of three-dimensional assemblies of magnetic nanoparticles for practical applications.

Acknowledgements

We gratefully acknowledge staff members of the Electron Microscopy Unit of Trinity College Dublin for their kind support during this work.

Notes and references

†*Synthesis of $\text{Fe}(\text{O}i\text{Bu})_2(\text{THF})_2$* : $\text{NaO}i\text{Bu}^1$ (0.68 g, 7.08 mmol) was added to FeBr_2 (0.76 g, 3.54 mmol) in THF (100 ml) at *ca.* 0°C . The mixture was stirred for 24 h at ambient temperature, filtered and the filtrate then concentrated to *ca.* 30 ml *in vacuo*. Grey crystals of $\text{Fe}(\text{O}i\text{Bu})_2(\text{THF})_2$ (0.64 g, 53%) were gradually deposited at 4°C . Further concentration of the remainder of the mother-liquor (*ca.* 15 ml) *in vacuo* and cooling to -22°C afforded an additional crop of $\text{Fe}(\text{O}i\text{Bu})_2(\text{THF})_2$ (0.31 g, 78% in total).

- 1 R. F. Service, *Science*, 2000, **287**, 1902.
- 2 F. E. Krus, H. Fissan and A. Peled, *J. Aerosol Sci.*, 1998, **29**, 511.
- 3 M. E. McHenry and D. E. Laughlin, *Acta Mater.*, 2000, **48**, 223.
- 4 J. Roger, J. N. Pons, R. Massart, A. Halbreich and J. C. Bacri, *Eur. Phys. J. Appl. Phys.*, 1999, **5**, 321.
- 5 C. Bergemann, D. Muller-Schulte, J. Oster, L. Brassard and A. S. Lubbe, *J. Magn. Magn. Mater.*, 1999, **194**, 45.
- 6 A. Jordan, R. Scholz, P. Wust, H. Fahling and R. Felix, *J. Magn. Magn. Mater.*, 1999, **201**, 413.
- 7 E. Matijevic, *Chem. Mater.*, 1993, **5**, 412.
- 8 (a) G. A. Held, G. Grinstein, H. Doyle, S. H. Sun and C. B. Murray, *Phys. Rev. B*, 2001, **6401**, 2408; (b) S. H. Sun, E. E. Fullerton, D. Weller and C. B. Murray, *IEEE Trans. Magn.*, 2001, **37**, 1239; (c) S. Mann and J. P. Hannington, *J. Colloid Interface Sci.*, 1988, 326; (d) M. Breulmann, H. Colfen, H.-P. Hentze, M. Antonietti, D. Walsh and S. Mann, *Adv. Mater.*, 1998, **10**, 237.
- 9 J. Rockenberger, E. C. Scher and A. P. Alivisatos, *J. Am. Chem. Soc.*, 1999, **121**, 11595.
- 10 (a) X. Cao, Yu. Koltypin, G. Katabi, R. Prozorov, I. Felner and A. Gedanken, *J. Mater. Chem.*, 1997, **7**, 1007; (b) R. Vijayakumar, Yu. Koltypin, I. Felner and A. Gedanken, *Mater. Sci. Eng. A*, 2000, **286**, 101; (c) R. V. Kumar, Yu. Yu. Koltypin, Y. S. Cohen, Y. Cohen, D. Aurbach, O. Palchik, I. Felner and A. Gedanken, *J. Mater. Chem.*, 2000, **10**, 1125.
- 11 A. R. West, *Solid State Chemistry and Its Applications*, John Wiley & Sons, London, 1984, p. 174.
- 12 J. M. D. Coey, A. H. Morrish and G. A. Sawatzky, *J. Physique*, 1971, **32**, C1-271.
- 13 S. R. Elliott, *Physics of Amorphous Materials*, Longman, London, New York, 1984, p. 350.